

## **Bottle Lab Instructions:**

**Step 1:** Write down the name of the chemical from the bottle. (Bottle # 9 in row 9)

**Step 2:** Write the cation. Include the symbol and charge. *Remember:* the charge comes from the group number. *Transitional metals will have the charge on the bottle already.*

**Step 3:** Write the number of valence electrons for the cation. *Remember:* Group # = valence electrons.

**Step 4:** Write the symbol and charge for the anion. *Remember:* The charge is equal to how many electrons it would need to gain to get to eight.

**Step 5:** Write the number of valence electrons for the anion. *Remember:* the charge comes from the group number.

**Step 6:** Write the formula for the compound. *Remember:* Use the criss-cross method.

**Step 7:** Determine the formula mass for each compound. *Note:* Add the atomic mass for each of the atoms in the element.